Graphitic carbon nitride (g-C₃N₄), a metal-free semiconductor with a band gap of 2.7 eV, has received considerable attention owing to its fascinating photocatalytic performances under visible-light. g-C₃N₄ exhibits high thermal and chemical stability and non-toxicity such that it has been considered as the most promising photocatalyst for environmental improvement and energy conservation. Hence, it is of great importance to obtain high-quality g-C₃N₄ and gain a clear understanding of its optical properties. Herein, we report a high-yield synthesis of g-C₃N₄ products via heating of high vacuum-sealed melamine powder in an ampoule at temperatures between 450 and 650 °C. Using transmission electron microscopy (TEM), scanning transmission electron microscopy (STEM), electron energy loss spectroscopy (EELS), thermogravimetric analysis (TGA), X-ray diffraction (XRD), and X-ray photoelectron spectroscopy (XPS), the chemical composition and crystallization of the as-produced g-C₃N₄ are demonstrated. A systematic optical study of g-C₃N₄ is carried out with several approaches. The optical phonon behavior of g-C₃N₄ is revealed by infrared and Raman spectroscopy, and the emission properties of g-C₃N₄ are investigated using photoluminescence (PL) spectroscopy, while the photocatalytic properties are explored by the photodegradation experiment.

Introduction

Graphitic carbon nitride (g-C₃N₄), one of the carbon nitride allotropes, has a graphene-like layered structure composed of heptazine units and the bridge amino groups. With a bandgap of 2.7 eV, g-C₃N₄ exhibits outstanding photocatalytic performance in various photochemical reactions, such as in photodegradation and photocatalytic water splitting under visible light. Unlike traditional organic semiconductors, g-C₃N₄ has excellent thermal and chemical stability, and it is also pollution-free, earth-abundant, inexpensive and facile to produce. Therefore, g-C₃N₄ is a very competitive candidate for depollution and solar energy development. The high-quality preparation and optical properties of g-C₃N₄ are thus of sustained interest. Great efforts were made to obtain high quality g-C₃N₄ in previous literature. High-yield synthesis of nano-sized g-C₃N₄ from bulk g-C₃N₄ has been reported recently as well. However, the most commonly used approach for the bulk g-C₃N₄ preparation is still by heating the reagent in air, such as cyanamide, dicyandiamide, or melamine. This one-step method is simple and easy to control; however, it suffers from a very low yield (∼6%) because the polymerization temperature of g-C₃N₄ is higher than the sublimation point of the reagent, causing the reagent loss before the reaction occurs. On the other hand, air contains many gas components such as oxygen and nitrogen, so there may be oxidation or other unexpected reactions during the heating process. To solve the problem on purity and yield, we choose to synthesize g-C₃N₄ in a vacuum-sealed environment. In this paper, all reagents are sealed inside an evacuated ampoule at a pressure of 3–10 mbar and then heated to the targeted reaction temperature. Since the reagent is not able to escape from the sealed ampoule, a high yield production of g-C₃N₄ (∼61%) is achieved. Additionally, there is negligible contamination or oxidation of the g-C₃N₄.
products since the heating process is performed in a vacuum. With TEM, STEM- EELS, TGA, XRD and XPS, the composition and structure of the as-produced g-C3N4 are characterized to demonstrate the good quality g-C3N4 produced via our synthesis approach.

Optimizing the optical properties of g-C3N4 is also of great importance to achieve better performance in photocatalytic reactions. As revealed by the state-of-the-art optical studies of g-C3N4, different heating temperatures of carbon nitride may lead to the change in the chemical structure of the sample. A blue- or red-shift of the photoluminescence may also occur. Following these studies, we obtained a series of products at various heating temperatures and carried out optical characterization including infrared, Raman, and PL spectroscopy to reveal the phonon behavior and the chemical bonds of the g-C3N4 products. We also observed a tunable PL emission in the g-C3N4 products processed at different temperatures and attribute this phenomenon to the extension of the g-C3N4 network at increasing temperature. Moreover, a photocatalytic decomposition experiment using g-C3N4 products is demonstrated, corroborating the excellent photocatalytic properties of the g-C3N4 products.

**Experimental section**

**Materials synthesis**

50 mg of melamine powder were firstly sealed in a quartz ampoule with a diameter of 13 mm and a length of 6–8 cm, after which the ampoule was pumped down to 3–10 mbar. Then, the ampoule is placed in the central region of a quartz tube furnace and heated at a ramp-up rate of 8 °C min⁻¹ to the designated temperatures between 450 and 650 °C. The temperature is then held constant for 2 hours. The ampoule was taken out after cooling to room temperature and opened by using a glass cutter. The powders inside the ampoule were collected into a 2 ml centrifuge tube and ultrasonicated in deionized water before dispersion. After 10 min centrifugation at 14 650 rpm, the supernatant liquid was discarded. This washing process is repeated 5 times to remove the unreacted melamine. After washing, the powder is dried in an oven at 60 °C for 3–5 hours.

**Characterization methods**

**TEM characterization.** The g-C3N4 products are dispersed in isopropl alcohol by ultrasonication and then deposited on a lacey carbon TEM grid. The morphology of the g-C3N4 products was observed by transmission electron microscopy (JEOL JEM 1400) with an accelerating voltage of 100 kV.

**STEM-EELS characterization.** The STEM-EELS studies of the g-C3N4 products were conducted by using a FEI Titan microscope operated at 80 kV under a high angle annular dark field (HAADF).

**TGA measurements.** The weight percent of each composition of the g-C3N4 products was determined by TGA (TA Instruments Q500) with a sampling rate of 10 °C min⁻¹ at 800 °C in air.

**XRD measurements.** XRD of the g-C3N4 products was performed using a Bruker D8 powder X-ray diffractometer with Cu Kα radiation (1.54056 Å) and a sampling rate of 0.8 degrees per min.

**XPS measurements.** To study the chemical bond of the g-C3N4 products, XPS measurements were performed on a VG ESCALAB 220i-XL system using a monochromatic Al Kα source. The pass energy of the analyzer was set at 10 eV to have a high measurement resolution. The XPS binding energy scale was calibrated with pure Au, Ag, and Cu standard samples by setting the Au 4f7/2, Ag 3d5/2, and Cu 2p3/2 peaks at binding energies of 83.98 ± 0.02 eV, 368.26 ± 0.02 eV, and 932.67 ± 0.02 eV, respectively. The binding energy calibrations were made against the C 1s peak to eliminate the charging of the sample during analysis.

**Optical property studies**

**Fourier transform infrared spectroscopy (FTIR).** FTIR (Perkin Elmer Frontier) is used to determine the chemical structure and the optical properties of the products by scanning from 600 cm⁻¹ to 4000 cm⁻¹ 16 times.

**Raman spectroscopy.** Raman scattering spectroscopy measurements were carried out with 785 nm excitation using a micro-Raman spectrometer (Horiba-JY T64000). The laser power is about 0.3 mW and the measurements were conducted in triple mode with 1800 g mm⁻¹ grating.

**Photoluminescence spectroscopy.** The photoluminescence properties were investigated using a micro-Raman spectrometer (Horiba-JY T64000) in the backscattering configuration excited using a He Cd laser (325 nm) with a power intensity of 0.3 mW (single mode, 600 g mm⁻¹ grating). The liquid nitrogen continuous flow cryostat (Cryo Industry of America, USA) was used to provide a continuous temperature variation from 77 to 300 K.

**Photocatalytic degradation experiment**

10 mg l⁻¹ methylene blue and 1 g l⁻¹ g-C3N4 products were used in the experiment and mixed at a stirring rate of 800 rpm in the dark for 30 min. Then the mixture was moved to the solar simulator for photocatalytic experiment at the same stirring rate. After each designated interval time (5 min, 5 min, 10 min, 10 min, 20 min, 20 min, 40 min), 1 mL sample is collected into a 2 ml centrifuge tube and centrifuged for 10 min at the rate of 14 650 rpm. Then, the UV-VIS absorption spectra of the supernatant liquid were recorded using a Lambda 950 spectrophotometer at room temperature.

**Results and discussion**

**Preparation and characterization of high-yield g-C3N4 products**

The formation of a highly crystalline carbon nitride usually follows the self-condensation process of cyanamide, dicyandiamide, or melamine during heating. In our case, we use...
melamine as the reagent in a sealed quartz ampoule. The vacuum level inside the ampoule is controlled under 3–10 mbar. With a heating ramp-up rate of 8 degree centigrade per minute, the ampoules were heated at different temperatures (450, 500, 550, 600, and 650 °C) for 2 hours.

During the heating process, the melamine will first form melem (an intermediate product during self-condensation of melamine) which is then polymerized into a g-C3N4 network, accompanied by a color change from white (melamine, melem) to yellow (g-C3N4).25 This phenomenon was also observed in our g-C3N4 products heated at different temperatures. Fig. 1(a) shows the photograph of all the products synthesized at different temperatures. The color of the products changes from white (450 °C) to pale yellow (500 °C), yellow (550 °C), brown (600 °C), and to dark brown (650 °C), which indicates that the composition of the products is changed. For the white color sample, the composition should be melem in majority. The yellower the sample becomes, the more g-C3N4 it contains. When the preparation temperature is too high (600–650 °C), the sample turns into brown and dark brown, which may be due to the carbonization of the sample at high temperatures.

Fig. 1(b–f) display the TEM images of the carbon nitride products processed at different temperatures. The morphology of products obtained at 450 °C and 500 °C have a unique rod-like structure, which we suspect to be the hydrogen-bonded framework of melem.26 Beside the rods, there are also a few layered flakes with a width of several hundred nanometers in products obtained at 450 °C and 500 °C. When the preparation temperature is above 550 °C, the flakes became the main products and the average size also increased to several micrometers.

To figure out the composition of the rods and flakes, STEM-HAADF and STEM-EELS have been conducted. Fig. 1(g) shows the STEM-HAADF image of a flake obtained at 600 °C, and Fig. 1(h and i) show the EELS chemical composition maps for C and N obtained from the flake. Chemical composition maps suggest a fairly homogeneous distribution of C and N throughout the flake. In general, the carbon nitride flake has a relative C composition of ~45 ± 1 at% and a relative N composition of ~55 ± 1 at%, which is almost equal to the nominal compositions of the C3N4 phase (C: 42.9%, N: 57.1%). Meanwhile, the EELS result from the bulk of the rod (Fig. S1†) also shows homogeneous distribution of C (~41 ± 0.5 at%) and N (~59 ± 0.5 at%). In this latest case, the C/N ratio is smaller than the nominal compositions of the C3N4 phase, suggesting a composition in good agreement with polymeric melem (C: 37.5–42.9%, N: 62.5–57.1%).

To confirm the content of the g-C3N4 products, we have conducted further characterizations. Thermogravimetric analysis (TGA) was performed to find the weight percent of each composition of the heated products. All the samples were heated from room temperature to 700 °C at a rate of 10 °C min⁻¹ in air, and the TGA curves are shown in Fig. 2(a). For the samples obtained at 450 °C and 500 °C, the weight loss at 100 °C for bound water (5–10%), 430 °C for melem (9–14%) and 550 °C for g-C3N4 (74–77%) was easily identified by their different temperatures of combustion or evaporation. Meanwhile, the samples obtained at 550–650 °C show a high purity of the g-C3N4 level (82–94%), which is in good agreement with the STEM-EELS results. We also observed the weight increase after 640 °C for the samples obtained at 550 °C–660 °C, and we suspect it might be due to the oxidation of the impurities. Considering the loss from the unreacted melamine, melem and impurities, we obtained 30.5 mg g-C3N4 from 50.0 mg melamine, so the highest yield of g-C3N4 can reach up to 61% when the heating temperature is 550 °C. Considering the theoretical yield of 73%, the reaction efficiency from melamine (C3N6H6) to g-C3N4 is almost 84%, this high yield has not been achieved in previous work.

Fig. 2(b) shows the X-ray diffraction (XRD) patterns of both melamine and the g-C3N4 products. As the preparation temperature increases, the melamine index peak was gradually eliminated. For all the g-C3N4 products, a strong new peak at 27.4° appeared, representing the typical stacking structure of g-C3N4 layers in the (002) direction.27–29 And for the samples obtained at 550 °C–650 °C, there are also few weaker XRD peaks at 13.0°, 44.5° and 57.5°, which represent the (100), (201) and (004) directions in the g-C3N4 crystal lattice separately.30 In the case of samples obtained at 450 °C and 500 °C, there is another strong peak at 11.0°, which is the character-
istic peak of the hydrogen-bonded framework of melem.\textsuperscript{26} This hydrogen-bonded framework is formed by the rearrangement of the melem molecules with the appearance of water, and its destruction can be achieved when the sample is heated above 100 °C (i.e., the dehydration temperature of crystal water); consequently, the characteristic peak at 11.0° then disappears (ESI, Fig. S2†).

X-ray photoelectron spectroscopy (XPS) was conducted to determine the chemical bonding of the g-C\textsubscript{3}N\textsubscript{4} products. In Fig. 2(c) and (d), the high-resolution C 1s and N 1s spectra of both melamine and the g-C\textsubscript{3}N\textsubscript{4} products have been collected and then deconvolved into several Gaussian peaks. In the case of melamine, the C1 peak at 287 eV represents the sp\textsuperscript{2} C atoms in the aromatic ring attached to the –NH\textsubscript{2} group.\textsuperscript{31} After the self-condensation of melamine during annealing, the C1 peak disappeared because of deamination. Instead, the C2 peak originated from sp\textsuperscript{2} C atoms bonded to N inside the aromatic structure, appeared at 288.5 eV.\textsuperscript{32,33} Remarkably, when the heating temperature is above 600 °C, a C3 peak appears at 291 eV. Considering the carbonization of the sample at high temperatures, this peak probably comes from the carbon π-π* transition.\textsuperscript{34} The N 1s spectra also show similar results. For melamine, the strong N2 peak at 399 eV represents the sp\textsuperscript{2} N atoms involved in triazine rings.\textsuperscript{35} Meanwhile, the N4 peak at 405 eV is contributed by the sp\textsuperscript{3} N atoms in the amino groups.

Fourier transform infrared spectroscopy (FTIR) and Raman studies of g-C\textsubscript{3}N\textsubscript{4} products

To investigate the change in the chemical structure during thermal condensation and to gain insight into the optical properties of the g-C\textsubscript{3}N\textsubscript{4} products, Fourier transform infrared spectroscopy (FTIR) and Raman spectroscopy were conducted on the samples. Fig. 3(a) shows the FTIR spectra of both melamine and the g-C\textsubscript{3}N\textsubscript{4} products. For melamine, the peaks at 3472, 3417, and 3325 cm\textsuperscript{-1} are attributed to stretching and deformation modes of –NH\textsubscript{2} groups.\textsuperscript{39} These peaks gradually decreased when the preparation temperature increases, indicating the process of deamination. Meanwhile, the peak at 807 cm\textsuperscript{-1}, which corresponds to the breathing mode of the triazine ring, remains.\textsuperscript{40} There is also a bunch of peaks appearing in the 1000–1700 cm\textsuperscript{-1} region, which are related to the
Emission properties of g-C3N4 products

To understand the emission properties of our g-C3N4 products, we further studied their photoluminescence (PL) properties. Fig. 4(a) shows the photographs of the multicolor-emission from melamine and the g-C3N4 products at the same concentration under UV light illumination (365 nm). We can clearly observe the color variation from blue-violet to green. Meanwhile, the intensity of the emission was enhanced at first and quenched later, indicating that high purity of g-C3N4 results in high PL intensity.

The steady-state and transient PL emission spectra have been excited by using a He–Cd laser (325 nm). The normalized photoluminescence spectra in Fig. 4(b) further prove the red shift of the luminescence center when the processing temperature increases from 450 °C to 650 °C. For the melam-rich g-C3N4 sample (450 °C), the center of PL spectra is at 369 nm; for the g-C3N4 sample (550 °C), the center of PL spectra is at 430 nm; for the carbonized g-C3N4 sample (650 °C), the center of PL spectra is at 455 nm and the peak also becomes broader because of impurities.

Gaussian fitting of the PL peaks helps us obtain a clear understanding of the nature and origin of excitons in the g-C3N4 sample. Three major emission centers have been demonstrated in the fitting and decomposition of the emission spectrum of the g-C3N4 samples (ESI Fig. S3†). Fig. 4(c) shows the lineshape analysis of the g-C3N4 sample prepared at 550 °C, which includes the emission center P1 (429 nm, 2.89 eV), P2 (451 nm, 2.75 eV) and P3 (484 nm, 2.56 eV). According to the previous PL study of g-C3N4, the bandgap states of g-C3N4 consist of a sp3 C–N σ band, sp2 C–N π band and the lone pair (LP) state of the bridge nitride atom (Fig. 4(d)), and the P1, P2 and P3 origin from the 3 different pathways of transitions: π*–π, σ*–LP and π*–LP respectively. Fig. 4(e) shows the red shift of P1, P2 and P3 with the preparation temperature increase. This red shift can be explained by the extension of the g-C3N4 network at higher temperatures. When more heptazine is connected by the amino group, the π states will hybridize into a broad state, causing the bandgap narrowing of the sp2 C–N clusters.20
Low temperature photoluminescence experiments have been carried out. Fig. 5(a–e) show the temperature-dependent PL spectra of the g-C₃N₄ sample obtained from 450 to 650 °C. As the temperature decreases, the features of the PL emission spectra are resolved clearer due to the weaker thermal vibration of atoms at low temperatures. Meanwhile, the peak position also shows the blue shift for about 5–10 nm when the temperature decreases from room temperature to 77 K (Fig. 5(f)). This tunable PL mechanism is mainly attributed to the size reduction of sp² C–N clusters when the temperature decreases. A smaller cluster size will result in a larger bandgap and higher probability of the direct transition between the π and π* bands. In addition, the LP valence band shifts in the direction of the π band at low temperatures, so the separation between the LP valence and π* bands will increase and cause a blue shift in the PL emission center. The photoluminescence of carbon nitride in deionized water has also been collected (ESI, Fig. S4†). With water involved, the PL spectra also show a slightly blue shift, which is probably due to the hydrogen-bonding existing in the carbon nitride network, and the LP valence band is affected.

**Photocatalytic degradation using g-C₃N₄ products**

To investigate the photocatalytic performance of our g-C₃N₄ products, photocatalytic degradation experiments have been performed. Methylene blue was used as the target for degradation and the degradation efficiencies were evaluated by measuring the peak value of the UV-Vis absorption spectra of methylene blue solution. Fig. 6(a) shows irradiation for different durations of 10 mg L⁻¹ methylene blue after photocatalytic degradation by the g-C₃N₄ sample (550 °C) under white light from a solar simulator. It can be clearly seen that the absorbance peak of methylene blue gradually decreased, and nearly went down to 0 after 110 min.

The degradation rates of all the g-C₃N₄ products (450–650 °C) are summarized in Fig. 6(b). With the increasing preparation temperature, the degradation rate of the g-C₃N₄ sample first increased, reached the highest point at 550 °C, and then finally decreased. Judging from the degradation rate, g-C₃N₄ samples with higher purity have higher photocatalytic activity, indicating that g-C₃N₄ has better oxidizability than the by-products of melem or carbonized g-C₃N₄. Moreover, com-

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**Fig. 5** (a–e) The normalized PL emission spectra of the g-C₃N₄ sample from room temperature (300 K) to liquid nitrogen temperature (77 K). The sample preparation temperature is labelled on the top-right panel. (d) The temperature-dependence of the PL peak position change for the g-C₃N₄ samples obtained at different temperatures.

**Fig. 6** (a) The absorption spectra of 10 mg L⁻¹ methylene blue after photocatalytic degradation by the g-C₃N₄ products synthesized at 550 °C at different time stages. (b) The normalized absorbance of methylene blue with respect to the elapsed time after the start of degradation reaction with the g-C₃N₄ products synthesized at different temperatures.
pared with previous literature, our g-C3N4 sample shows a similar photocatalytic performance to the normal air-heated g-C3N4 but with a much higher yield.3,46

Conclusions

In conclusion, we successfully prepared a high yield of g-C3N4 products using sealed thermal condensation of melamine and investigated the compositional change with the increase of growth temperature. The optical properties and the internal structure of the g-C3N4 products have been explored with both FTIR and Raman spectroscopy. Moreover, the PL properties of the g-C3N4 products synthesized at different temperatures have been studied. With the increase of synthesis temperature, the carbon nitride products exhibit tunable PL properties. The PL centers vary from blue-violet light region to green light region, which may indicate the optical band gap changes during synthesis. Furthermore, we studied the photocatalytic degradation properties of the g-C3N4 products, and concluded that the highest photocatalytic efficiency was obtained at ~550 °C growth temperature.

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Notes and references
